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18. Use the bond energies in Model 2 to calculate the energy that is released to form the two moles of answer. 19. The heat of reaction or change in enthalpy (ΔH) of a reaction is the net energy change in the

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A good answer should state that as the distance between a proton (nucleus) and electron increases, the attractive force decreases. This is followed by a quick class discussion to summarize the previous day's lesson and leads quickly into model 3 and 4 of the POGIL activity.

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Chemical Bonding POGIL Activity 7 - A Mixed Up Activity Answer Key # 1 & #2. Chemical Formula Compound Name Monatomic Ions (#1) Polyatomic Ions (#2) CaSO₄ 4 calcium sulfate calcium, Ca +2 sulfate, SO₄ -2 CaSO₃ 3 calcium sulfite calcium, Ca +2 sulfite, SO₃ -2 Na₃PO₄ 4 sodium phosphate sodium, Na +1 phosphate, PO₄ -3.

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Pogil Bond Energy Answer Key Use the bond energies in Model 2 to calculate the energy that is released to form the two moles of molecules in Step 2 of the reaction in Model 3. Include the proper sign and units on your answer.

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Calorimetry Pogil Answers - examred.com Equation for the calculation of heat is: $q = mC\Delta T$ q = heat released in Joules. m = mass of water in the calorimeter. C = specific heat capacity (not 0°C which is a temperature) ΔT = [final temperature - initial temperature] (absolute value) Key Questions: Calorimetry - Measurement of Heat Energy

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Chemical Bonding POGIL Activity 7 | Chemistry Quiz - Quizizz
4 POGIL™ Activities for High School Chemistry 14. Often we draw Lewis structures with 90° bond angles. Do any of the molecular shapes in Model 1 have 90° bond angles? 15. Which shape was used twice in Model 1? What do the two Lewis Structures have in common that cause them to be the same shape? 16.

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Clz a. $\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow$ b. $\text{Fe(s)} + 2\text{HCl(aq)} \rightarrow$ c. $\text{Cu(s)} + \text{HCl(aq)}$
13. In old westerns, we often see a desperate prospector get very excited when he sees a vein of gold in the wall of his mine. It is true that gold, silver, and copper are typically found in nature as met- als.

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