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# **Strong Versus Weak Acids Pogil Answer Key Knutke**

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S - Strong versus Weak Acids What  
... 022 Gas Variables Worksheet**

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## **POGIL Chemistry Activities STRONG ACIDS vs. WEAK ACIDS**

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Strong versus Weak Acids 1 Strong  
versus Weak Acids What makes a strong  
acid strong? Why? Acids are substances

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that surround us in our everyday life. The uses of acids range from providing essential nutrients for our bodies to dissolving metals. Some acids are safe to handle with our bare hands or even use in food preparation.

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versus Weak Acids What makes a strong  
acid strong? Why? Acids are substances  
that surround us in our everyday life.  
The uses of acids range from providing



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essential nutrients for our bodies to dissolving metals. Some acids are safe to handle with our bare hands or even use in food preparation. Other acids will severely burn human skin.

### Introduction to Chemistry Atomic and Electron Structure

Lewis Theory of acids and bases-POGIL

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2; Acids and Bases supplement Acid Strength-POGIL 3; Quiz 1: Wednesday Nov 22nd Objectives: 1- Define three different theories of acids and bases 2- Identify and distinguish Arrhenius, Bronsted and Lewis Acids and Bases ... 7- State six strong Acids 8- State Organic acids as weak Acids.

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### List of Common Strong and Weak Acids

Strong versus Weak Acids 3 5. Based on the data in Model 1 and the table in Question 3, describe the relationship between:

- the percent ionization of the acid and the conductivity of the solution.
- the conductivity of the solution and the strength of the electrolyte (acid strength).

6. Consider the conductivity

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data shown in Model 1 and the ionization data in Question 3.

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Strong versus Weak Acids 3 5. Based on the data in Model 1 and the table in Question 3, describe the relationship between: a. the percent ionization of the acid and the conductivity of the solution.

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b. the conductivity of the solution and the strength of the electrolyte (acid strength). 6. Consider the conductivity data shown in Model 1 and the ionization data in Question 3.

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## Strong versus Weak Acids - Science Done Wright

Some weak bases are the conjugate bases of weak acids. These are generally the anions produced by the acid's hydrolysis. Like molecular bases, these conjugate bases have base hydrolysis



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equilibria with associated  $K_b$  values. The relationship between  $K_a$  and  $K_b$  for a conjugate acid-base pair allows us to calculate the pH of a solution of either species.

Chem 116 POGIL Worksheet - Week 10  
Common Ion Effect and ...  
Strong vs Weak; Sources; Strong Acids

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and Bases. A strong acid or base is one that is completely ionized in a solution.

Strong Acids: - Hydrochloric Acid ( HCl ) -  
Hydrobromic Acid ( HBr ) - Sulfuric Acid ( H<sub>2</sub>SO<sub>4</sub> ) -  
Perchloric Acid ( HClO<sub>4</sub> ) -  
Nitric Acid ( HNO<sub>3</sub> ) - Hydroiodic Acid ( HI )

## STRONG AND WEAK ACIDS

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## Answer Key Knutke Weak Acid and Base Equilibria

Compared to a weak Arrhenius acid, a strong Arrhenius acid. a. is more soluble in water b. is a better oxidizing agent c. is more highly ionized on water solution d. has more available protons per molecule e. has stronger bonds between hydrogen and oxygen atoms.

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## 33 Strong vs Weak Acids-S - Strong versus Weak Acids What ...

With strong acids this is easy.

Hydrochloric acid is a strong acid - virtually 100% ionised. Each mole of HCl reacts with the water to give 1 mole of hydrogen ions and 1 mole of chloride ions. That means that if the concentration of the acid is 0.1 mol

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dm<sup>-3</sup>, then the concentration of hydrogen ions is also 0.1 mol dm<sup>-3</sup>.

### 022 Gas Variables Worksheet (1).pdf - Gas Variables How ...

Adding small amounts of a strong acid (e.g., HCl) will neutralize an equivalent amount of the conjugate base and form an equivalent amount of the weak acid.

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Conversely,

## POGIL Chemistry Activities

™ POGIL Activities for High School  
Chemistry Chemical Reactions and  
Stoichiometry Types of Chemical  
Reactions . . . . .  
. . .153 .



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Strong acids completely dissociate into their ions in water, while weak acids only partially dissociate. There are only a few (6) strong acids, so many people choose to memorize them. All the other acids are weak. The strong acids are hydrochloric acid, nitric acid, sulfuric acid, hydrobromic acid, hydroiodic acid,

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perchloric acid, and chloric acid.

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